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to answer the following: Mass of the empty crucible 4.00 g Mass of the crucible plus hydrate sample 4.50 g Mass of the system after heating 4.32 g Mass of the system after a second heating 4.32 g 36% a. Determine the mass percentage of water in the original sample. MODERN CHEMISTRY CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS59

### **7 Chemical Formulas and Chemical Compounds**

Your alarm goes off and, after hitting "snooze" once or twice, you pry yourself out of bed Modern chemistry answer key chapter 2. You make a cup of coffee to help you get going, and then you shower, get dressed, eat breakfast, and check your phone for messages.

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Explain your answer. The forward reaction is favored. The weaker acid and weaker base are produced in the forward reaction. HCO<sub>3</sub><sup>-</sup> competes more strongly with H<sub>2</sub>O to donate a proton, and OH<sup>-</sup> competes more strongly with CO<sub>3</sub><sup>2-</sup> to acquire a proton, causing the forward reaction to be favored. MODERN CHEMISTRY ACIDS AND BASES 121

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## **2 Measurements and Calculations**

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